Bases, and Solutions " Rewiccribing Acids and scteribing Main Ideas ete the following table.  acteristic Acid Acid acteristic Acid acteristic Acides it taste like?  does it react with the does it react react react react react react carbonates?  It color does it litmus?  S. carbonates:  It the following questions in the sper the following questions in the sper the following actions in the sper the following acti	<u></u>	-	10.	9 A m	T	s 171	2 7 7	\$ \$	0	Acids, Des Unde
Is and Bases (A). HOY-FOP  eas  Acid  Acid  Base  1.  2.  3.  4.  5.  5.  6.  5.  8.  what may betance?  betance?	12. Why do you think bas	11. What is an indicator?		Building Vocabulary  Answer the following questic  9. What does corrosive m	What color does it turn litmus?	How does it react with carbonates?	How does it react with the metals magnesium, zinc, and iron?	When found in foods, what does it taste like?	Characteristic	Acids, Bases, and Solution Describing Acid Understanding Main Id Complete the following table.
Base  Base  6.  6.  6.  1.  1.  1.  1.  1.  1.  1.	ses are often described as t		rith a metal to produce hydbstance?	ms in the spaces provided. ean?	7.	Çī		<del>jun</del>	Acid	Review and Reinfo
	ne "opposites" of acids?		lrogen gas, what may			6.	ję.	12	Base	aps. 404-409

## is and Bases in Solution Pys. 700-715

## rstanding Main Ideas

ste the concept map shown below and answer the following questions on a e sheet of paper.

	la water they form
Substances with a pH in this range are 2.	Substances with a pH in this range are

- hat is the difference between a strong acid and a weak acid?
- hat is the difference between a strong base and a weak base?
- hich solution has a greater concentration of hydrogen ions (H $^+$ ), a plution with a pH of 3 or one with a pH of 7? Explain.
- hat are the products formed when a base reacts with an acid?
- That is the pH of a neutral solution?

## ling Vocabulary

each term with its definition by writing the letter of the correct symbol or thon on the line beside the term in the left column.

- 10. hydroxide ion
- ionic compound that can form from the reaction of an acid with a base

11. pH scale

- reaction between an acid and a base
- 12. hydrogen ion
- series of numbers that indicates the concentration of hydrogen ions in solution
- 13. neutralization OH-

14. salt