

**Acids, Bases, and Solutions \* Review and Reinforce**

**Describing Acids and Bases**

pgs. 704 - 709

**Understanding Main Ideas**

Complete the following table.

Characteristic	Acid	Base
When found in foods, what does it taste like?	1.	2.
How does it react with the metals magnesium, zinc, and iron?	3.	4.
How does it react with carbonates?	5.	6.
What color does it turn litmus?	7.	8.

**Building Vocabulary**

Answer the following questions in the spaces provided.

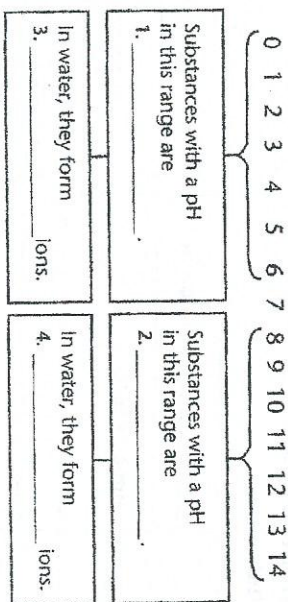
- What does *corrosive* mean?
- If a substance reacts with a metal to produce hydrogen gas, what may you infer about the substance?
- What is an indicator?
- Why do you think bases are often described as the "opposites" of acids?

**Acids and Bases in Solution**

pgs. 710 - 715

**Understanding Main Ideas**

Complete the concept map shown below and answer the following questions on a separate sheet of paper.



- What is the difference between a strong acid and a weak acid?
- What is the difference between a strong base and a weak base?
- Which solution has a greater concentration of hydrogen ions ( $H^+$ ), a solution with a pH of 5 or one with a pH of 7? Explain.
- What are the products formed when a base reacts with an acid?
- What is the pH of a neutral solution?

**Building Vocabulary**

Match each term with its definition by writing the letter of the correct symbol or definition on the line beside the term in the left column.

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|--------------------|--|
| 10. hydroxide ion  | a. ionic compound that can form from the reaction of an acid with a base           |
| 11. pH scale       | b. reaction between an acid and a base   |
| 12. hydrogen ion   | c. series of numbers that indicates the concentration of hydrogen ions in solution |
| 13. neutralization | d. $H^+$   |
| 14. salt           | e. $OH^-$  |